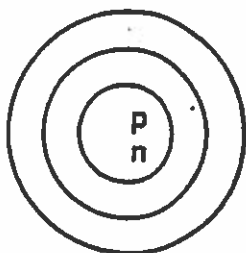
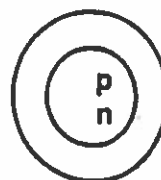


I. Bohr Model: Use the information provided for each element to complete the diagram. Place the correct number of electrons in their proper shell. Place the correct numbers in the nucleus to indicate the number of protons and the number of neutrons. Show your work.

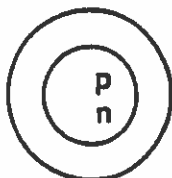
1. Sulfur: atomic number 16  
atomic mass 32



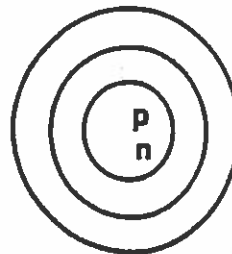
2. Beryllium: atomic number 4  
atomic mass 9



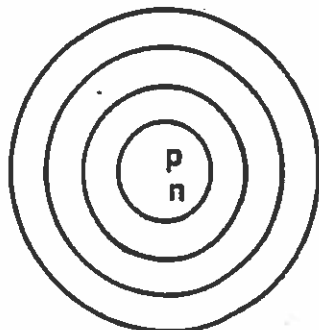
3. Nitrogen: atomic number 7  
atomic mass 14



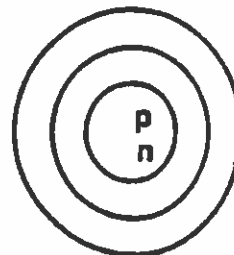
4. Sodium: atomic number 11  
atomic mass 23



5. Potassium: atomic number 19  
atomic mass 39

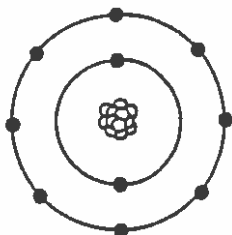


6. Argon: atomic number 18  
atomic mass 40

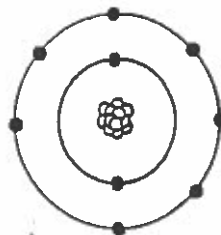


## II. Making Predictions: Applying Main Ideas

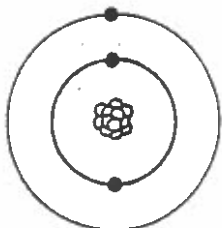
Predict whether each atom in the following diagrams will gain electrons, lose electrons, or be inert and by how many electrons. Next, determine the chemical symbols for which the diagram is illustrating.



1. \_\_\_\_\_



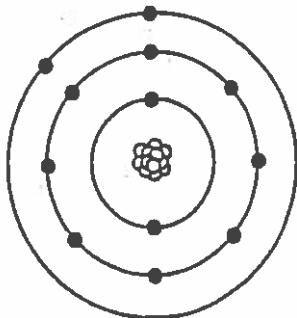
2. \_\_\_\_\_



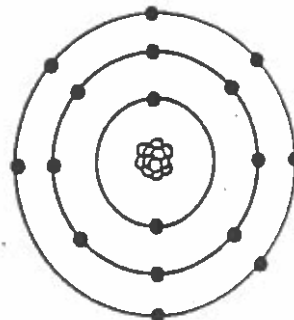
3. \_\_\_\_\_



4. \_\_\_\_\_



5. \_\_\_\_\_



6. \_\_\_\_\_